

✓ The reaction of some 2-substituted derivatives of 4-aminopyridine with nitrous acid. IV. 2-Methoxy-4-amino-pyridine. Tadeusz Talik and Edwin Płazek (Politechnika, Wrocław, Poland). *Kośniki Chem.* 33, T3-13-8 (1959) (German summary); cf. *CA* 50, 12045f; 53, 18954d. —4-Amino-2-methoxypyridine was able to be diazotized. 4-Hydroxy-2-methoxypyridine was able to be diazotized. 4-Hydroxy-2-methoxypyridine (m. 135°, yield 59.8%), 4-iodo- (b.p. 106°, 31.7%), 4-chloro- (m. 26°, b. 177-8°, 30.6%), 4-bromo- (b.p. 87°, 26.4%), 4-thiocyanato- (m. 62-3°, 11.2%), and 4-cyano-2-methoxypyridine (I) (m. 95°, 32.4%) were obtained from the diazo-compd. Thus, substitution of a first-order group (no double or semipolar bond) in 2-position made the 4-amino group like a normal aromatic amine. Hydrolysis of I gave 2-hydroxyisonicotinic acid. A. Kręglewski

TALIK, Zofia; TALIK, Tadeusz

On the reaction of 2-halogen and 3-halogen-4-nitropyridine
N-oxides with phosphortrihalogenides. Rocznik chemii 36 no.3:417-423
'62.

1. Katedra Chemii Organicznej I. Politechnika, Wrocław, i
Zakład Chemii Organicznej Katedry Chemii, Wyższa Szkoła
Ekonomiczna, Wrocław.

S/081/62/000/021/022/069
B141/B101

AUTHORS: Talik, Tadeusz, Talik, Zofia

TITLE: Preparation of some derivatives of 4-nitro-pyridine N-oxide

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 21, 1962, 173, abstract
21Zh163 (Roczn. chem., v. 36, no. 3, 1962, 539-544 [Pol.;
summary in Ger.])

TEXT: The action of a mixture of $(CH_3CO)_2O$ and H_2O_2 on pyridine,
 α -picoline, 2,6-lutidine, 2-chloro-, 2-bromo-, 2-iodo-, 3-chloro-,
3-bromo-, and 3-iodo pyridine (Ia-i), followed by nitration, yields the
N-oxides of 4-nitro-I (IIa-i). The structure of IIg, i was confirmed by
conversion with aniline to the N-oxide of 3-anilino-4-nitro-pyridine (III).
30 g Ia dissolved in 150 ml $(CH_3CO)_2O$ 150 ml receives an addition of 30%
 H_2O_2 in portions, the solution is left to stand for 5-6 hrs at $20^{\circ}C$ and is
then kept for 30 hrs at $60-65^{\circ}C$, the excess of reagents is evaporated in
vacuo at $60^{\circ}C$, the residue is dissolved in 50 ml concentrated H_2SO_4 and

Card 1/2

Preparation of some derivatives ...

S/061/62/000/021/022/069
B141/B101

added in portions to a mixture of 50 ml concentrated H_2SO_4 , 20 ml 20% fuming sulfuric acid, and 120 ml concentrated HNO_3 (d 1.52); the mixture is kept for 90 min at $100^{\circ}C$, cooled, and poured onto ice, neutralized to pH = 4-5 with solid $(NH_4)_2CO_3$, whereupon IIa is separated with concentrated NH_4OH , yield 63%, m.p. $162^{\circ}C$ (from water). Likewise substance II is obtained (yield in %, m.p. in $^{\circ}C$). b, 76.5, 155; c, 76.5, 156; d (here and below, 150 ml 30% H_2O_2 is added and the mixture heated for further 30 hrs) 69.4, 154; e, 62.5, 145; f, 61.5, 140; g ($C_5H_3ClN_2O_3$) (separated by extraction with $CHCl_3$), 54.5, 115; h, 54.2, 152; i ($C_5H_3IN_2O$) (obtained analogously to IIa), 56.4, 201. From the solution of 1 g IIg, 3 g aniline, and 20 ml alcohol which is boiled for 3 hrs and cooled, III, $C_{11}H_9N_3O_3$, is obtained, yield 0.7 g, m.p. $197-198^{\circ}C$ (from alcohol). Analogously, III is obtained from 1 g IIi and 3 g aniline in 40 ml alcohol by boiling for 6 hrs, yield 0.6 g. [Abstracter's note: Complete translation.]

Card 2/2

S/081/62/000/021/023/069
B141/B101

AUTHORS: Talik Zofia, Talik Tadeusz

TITLE: Effect of persulfuric acid on some 4-aminopyridine derivatives

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 21, 1962, 173, abstract 21Zh164 (Roczn. chem., v. 36, no. 3, 1962, 545-548 [Pol.; summary in Ger.])

TEXT: When β -chloro-4-aminopyridine (I), β -bromo-4-aminopyridine (II), and 2-bromo-4-aminopyridine (III) are brought into reaction with $H_2S_2O_8$, the products are β -chloro-4-nitro-pyridine (IV), β -bromo-4-nitro-pyridine (V) and 2-bromo-4-nitro-pyridine (VI) respectively. Under analogous conditions, 2-chloro-, 2-iodo-, and β -iodo-4-aminopyridine do not react with $H_2S_2O_8$. To 30 ml 20% fuming sulfuric acid 20 ml H_2O_2 is added dropwise under cooling by a mixture of ice and salt, then 2.5 g I dissolved in 7.5 ml concentrated H_2SO_4 is introduced at $0^{\circ}C$, the mixture is left to stand for 24 hrs at $20^{\circ}C$, poured onto ice, and neutralized with solid

Card 1/2

Effect of persulfuric acid on ...

S/61/62/000/021/023/069
B141/B101

$(\text{NH}_4)_2\text{CO}_3$; the precipitate is dissolved in ether, and evaporation of the solvent produced the compound IV, $\text{C}_5\text{H}_4\text{ClN}_2\text{O}_2$, yield 71.2%, m.p. 25-26°C (from benzene). Analogously V, $\text{C}_6\text{H}_4\text{BrN}_2\text{O}_2$, is obtained from 2.5 g II, yield 65.2%, m.p. 66-67°C. When 1 g III is brought into reaction with 3 ml H_2SO_4 and $\text{H}_2\text{S}_2\text{O}_8$ (from 8 ml H_2O_2 and 12 ml 20% fuming sulfuric acid), VI is obtained, yield 76.7%, m.p. 62°C (from alcohol). [Abstracter's note: Complete translation.]

Card 2/2

WIECZOREK, Jan Sylwester; TALIK, Tadeusz

Preparation and some reactions of 3-bromo-4,5-diaminopyridine.
Rocznik chemii 36 no. 5: 967-970 '62.

1. Katedra Chemii Organicznej I, Politechnika, Wrocław, i
Katedra Chemii, Zakład Chemii Organicznej, Wyższa Szkoła
Ekonomiczna, Wrocław.

TALIK, Tadeusz

Certain reactions of 3-halogen-4-nitropyridine N-oxides. Roczn
chemii 36 no.10:1465-1475 '62.

1. Katedra Chemii Organicznej I Politechnika, Wroclaw, i Katedra
Chemii, Zaklad Chemii Organicznej, Wyzsza Szkola Ekonomiczna ,
Wroclaw.

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

POLSKA, Tadeusz

Relative reactivity of halogens in some 3-halogen derivatives of
pyridine and their N-oxides. Roczn. chemii 36 no.11:1563-1573
'61.

I. Katedra Chemii Organicznej I, Politechnika, Wrocław,
II. Katedra Chemii, Zakład Chemii Organicznej, Wyższa Szkoła
Ekonomiczna, Wrocław.

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

TALIK, Tadeusz

Some reactions of 3-halogen-4-nitropyridines. Rocznik chemii 37
no.1:69-74 '63.

1. Zaklad Chemii Organicznej, Wyższa Szkoła Ekonomiczna,
Wrocław.

TALIK, Tadeusz; TALIK, Zofia

Reaction of 2-hydroxy-4-aminopyridine with nitrous acid. Rocznika
chemii 37 no.1:75-80 '63.

1. Zaklad Chemii Organicznej, Wyższa Szkoła Ekonomiczna,
Wrocław.

TALIK, Tadeusz

Reaction of 3-bromo-4-nitropyridine N-oxide with amino acids.
Rocznik chemii 37 no.4:495-497 '63.

I. Zaklad Chemii Organicznej, Wyższa Szkoła Ekonomiczna, Wrocław.

TADKZ TULIK

POL.

"Syntheses of pyridine analogs of *p*-aminosalicylic acid." Tadeusz Tulk and Zofia Tulik. *Acta Polon. Pharm.* 11, 71-5(1954)(English summary). By the Kolbe-Schmidt method, modified by Marasse (Ger. 73,279 and 78,700), 2-amino-6-hydroxy-6(?) (I), and 3-amino-5-hydroxy-6(?)-pyridinecarboxylic acid (II) have been prepd. I is obtained in 78.5% yield by heating 2-amino-6-hydroxypyridine with anhyd. K_2CO_3 under a CO_2 atm. in an autoclave 3-4 hrs. at about 200°, treating the mixt. with hot H_2O , filtering, and pptg. the product with glacial AcOH; it is difficultly sol. in cold H_2O , alc., C_6H_6 , $CHCl_3$, and acetone, and decomp. 185°. By the same procedure (reaction time 7-8 hrs.) 3-amino-5-hydroxypyridine yields II, isolated as the Cu salt, which with H_2S soln. gives 57% free II, m. 204-5°. sol. in hot H_2O , insol. in alc. and C_6H_6 . Both acids are inactive against tubercle bacilli. Michael Dymock

3

BS

TALIK, Zofia; PLAZEK, Edwin

Preparation of chlorides of 2- and 4-pyridino sulfonic acids and
of 2- and 4-pyridonesulfonamides. Acta Poloniae pharm. 12 no.1:
5-12 1955.

1. I Katedra Chemii Organicznej Politechniki Wrocławskiej. Kierownik:
prof. dr E. Plazek.

(PIPERIDINES, preparation of
chloropiperidinesulfonic acids & chloropiperidinesulfonamides)
(SULFONAMIDES, preparation of,
chloropiperidinesulfonamides)

TALIK Z.

✓ **Synthesis of pyridinesulfohydroxamic acids.** Z. Talik and T. Talik [Polytech., Wrocław, Poland]. *Acta Polon. Pharm.*, 12, 213-22 (1955) (English summary).—2-, 3-, and 4-Pyridinesulfohydroxamic acids (I), (II), and (III), resp., prep'd. by the reaction of the corresponding sulfonyl chlorides with hydroxylamine showed only small tuberculostatic activity in experiments *in vitro*. Na (2 g.) is dissolved in 30 ml. abs. EtOH and added below 40° to 7 g. NH₂OH.HCl (IV) in 5 ml. water. To the filtered soln., 2-pyridinesulfonyl chloride [obtained from 4 g. mercaptopyridine (cf. *ibid.* 5-12)] is added in portions below 30° and the mixt. kept 10 min., the EtOH is distd. *in vacuo*, 5 ml. water added, the soln. heated with C at 90° and filtered, and the ppt. crystd. from water to give I, m. 129-30° (decomp.). II is obtained by adding 13.5 g. IV in 10 ml. of water to a soln. of 4.25 g. Na in 60 ml. abs. EtOH at 40°, filtering off the NaCl, adding 3-pyridinesulfonyl chloride to the filtrate, filtering off IV, evapg. the filtrate and recrystg. the residue from water; II, m. 149-51° (decomp.). III is obtained similarly (no m.p. is reported).

R. Dowbenko

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Malin, S.; Malink, ...

"Nitration of 2-chloro-4-aminopyridine."

p. 1139 (Roczniki Chemii) Vol. 31, no. 4, 1956
Warsaw, Poland

SU: Monthly Index of East European Accessions (LEAI) LC. Vol. 7, no. 4,
April 1958

14
4E3d
Z. J. J. (WB)

Synthesis of certain derivatives of thiocyanacetic acid hydrazide. Zofia Talik and Edwin Plaizek (Politechnika, Wrocław, Poland). Roczniki Chem. 33, 379-85 (1959) (German summary.)—Attempts to prep. $\text{NCSCH}_2\text{CONH}_2$ (I) were unsuccessful, but 3 stable *hydrazones* ($\text{NCSCH}_2\text{NH}_2$) (II) were obtained as described below. Into 3.8 g. $\text{PhCH}_2\text{NNH}_2$ in 10 ml. $\text{C}_2\text{H}_5\text{N}$ at -30° was dropped 3.4 g. ClCH_2COCl in 20 ml. Et_2O , the mixt. then poured into 400 ml. H_2O , and the solid filtered off and recrystd., from $\text{H}_2\text{O}-\text{EtOH}$ to yield 3.4 g. $\text{ClCH}_2\text{CONHNHC}_6\text{H}_4\text{CHR}$ (III) ($\text{R} = \text{Ph}$), m. 164° . Similarly were obtained the analogs III ($\text{R} = m\text{-O}_2\text{NC}_6\text{H}_4$), m. 207° , and III ($= \text{R} = o\text{-HO}_2\text{C}_6\text{H}_4$), m. $198-9^\circ$. III with KSCN in Me_2CO gave II ($\text{R} = \text{Ph}$), m. 251° (alc.), II ($\text{R} = m\text{-O}_2\text{NC}_6\text{H}_4$), m. 278° (AcOH); and II ($\text{R} = o\text{-HO}_2\text{C}_6\text{H}_4$), m. 295° (decompn.). IIa ($\text{R} = \text{Ph}$) and IIb ($\text{R} = m\text{-O}_2\text{NC}_6\text{H}_4$) ($\text{C}_6\text{H}_5\text{N}$) showed strong tuberculostatic action *in vitro*. Double hydrazides of chloro-, m. 161° , and thiocyanacetic acid, decomp. 200° , were also prep'd.

TALIK, Z.; PLAZEK, E.

Investigations on 2-chloro-3,5-dinitropyridine. I. Exchange reactions
of the halogen atoms. II. Experiments comparing the movability of the
halogen atom. III. Possibilities of using 3,5-dinitro-2-chloropyridine
in investigations on protein adsorption and amino acids. Bul chim
PAN 8 no.5:219-230 '60. (EEAI 10:9/10)

1. Katedra Chemii Organicznej I., Politechnika, Wroclaw. Presented
by T. Urbanski.

(Chloro-dinitropyridine) (Halogens)
(Amino acids) (Proteins)

TALIK, ZOFIA
PLAŻEK, EDWIN

Distr: 4E3d

✓ 2-Chloro-3,5-dinitropyridine. Zofia Talik and Edwin
Plażek (Tech. School, Wrocław, Poland). Roczniki Chem.
34, 185-76 (1980). Several expts. concerning the unusually
high mobility of the Cl atom in 2-chloro-3,5-dinitropyridine
were carried out, in particular the substitution of the former
by CH₃, OMe, OEt, OCH₂CH₃O, OPh, SPh, SH, —SS—
—S—, SCN, and —CN groups. Some expts. were carried
out for this purpose on 2-bromo-3,5-dinitropyridine. The
high mobility of the halogen atom in these compounds was
confirmed. M. Treanor

4
2-Jag(CN)(may)

TALIK, Zofia

Research on 2-chloro-3,t-dinitropyridine. II. Reactions with animes.
Rocznik chemii 34 no.2:465-474 '60. (EEAI 10:1)

1. Katedra Chemii Organicznej I Politechniki, Warszawa.
(Chlorodinitropyridine) (Amines)

TALIK, Zofia

Investigations on 2-chloro-3,5-dinitropyridine. III. The relative reactivity of halogen. Rocznik chemii 34 no.3/4:917-924 '60.
(EEAI 10:3)

1. Katedra Chemii Organicznej I Politechniki, Wrocław.
(Halogens) (Chlorodinitropyridine)

L0192
S/081/62/000/013/015/054
B158/B144

5/14/01

AUTHOR:

Talik, Z.

TITLE:

Substitution of a halogen and a nitro-group in N-oxides of 2-halogen-4-nitropyridine

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 13, 1962, 245, abstract 13Zh217 (Bull. Acad. polon. sci. Ser. sci. chim., v. 9, no. 9, 1961, 561-565)

TEXT: For the study of the mobility of a halogen and the NO_2 group in N-oxides of 2-halogen-4-nitropyridine (I 4-nitropyridine), the latter were reacted with NaOCH_3 and NaOC_2H_5 at 20°C . The NO_2 group being easily substituted on the alkoxyl, the halogen was not very active under these conditions. In a reaction of 1 mole of the starting substance with 2 moles of NaOCH_3 under severe conditions (boiling), the halogen may be replaced by an alkoxyl. In a reaction with Fe(OE)_2 or PCl_3 (heating), the N-oxides of 2-halogen-4-alkoxypyridines lose the O atom. Under the

Card 1/4

S/081/62/000/013/015/054
B158/B144

Substitution of a halogen and a ...

effect of $\text{NH}(\text{CH}_3)_2$ or $\text{NH}(\text{C}_2\text{H}_5)_2$, N-oxides of 2-halogen-I are easily converted to N-oxides of dialkylamino-I, which by reacting with PCl_3 in CHCl_3 are converted to 2-dialkylamino-I; the latter are converted to 2-dialkylamino-4-aminopyridines (II 4-aminopyridine) by strong reducing agents. When N-oxides of 2-halogen-I are reacted with $\text{NH}_2\text{NH}_2 \cdot \text{H}_2\text{O}$ in alcohol, various mixtures are formed; only in one case is it possible to separate the N-oxide of 2-iodine-II. By the methods described, the following were obtained (product obtained, yield %, m. p. in $^{\circ}\text{C}$, m. p. in $^{\circ}\text{C}$ of picrate (supplements in brackets) are given): N-oxide of 2-chloro-4-methoxy pyridine (III 4-methoxy pyridine), 84.1, 82 (from benzene-dimine), 141; N-oxide of 2-bromo-III, 85.4, 55, (from benzene), 145; N-oxide of 2-iodo-III, 88.6, 89 (from water), 142; 2-chloro-III, 83.4 (b. p. 229-230 $^{\circ}\text{C}$), -, 168; 2-bromo-III, 74.6 (b. p. 121 $^{\circ}\text{C}/10 \text{ mm}$), -, 115; 2-iodo-III, 74.6, 35 (from benzene), 154; N-oxide of 2,4-dimethoxy pyridine, 72.1, 85 (from benzene), -; 2,4-dimethoxy pyridine, 62.6 (b. p. 200-201 $^{\circ}\text{C}$), -, 159; 2-chloro-4-ethoxy pyridine, 66.4, 55-57 (from

Card 2/4

S/081/62/000/013/015/054
B158/B144

Substitution of a halogen and a ...

benzine), 139; 2-bromo-4-ethoxy pyridine, 49.3, 38 (from benzine), 115; 2-iodo-4-ethoxy pyridine, 69.2, (b. p. 145°C/11 mm) -, 133; N-oxide of 2-dimethylamino-I, 74.8, 126 (from benzene-benzene), -, N-oxide of 2-diethylamino-I, 62.3, 90 (from benzene-benzene), -, 2-dimethylamino-I, 67.7, 99 (from benzene-benzene), -, 2-diethylamino-I, 75.8, 33 (from benzene), -, 2-dimethylamino-II, 62, 151 (from water, 216 (various); benzine), 76.7, 117 (from aqueous alcohol), 172; N-oxide of 2-diethylamino-II, 40.6, 110 (various; from water), 189. The solubilities of the enumerated compounds in water, alcohol, C₆H₆ and CHCl₃, are given.

N-oxides of 2-halogen-III are unstable in air, at 0°C they can be preserved for a long period, at 20°C they rapidly decompose, and at 100°C they ignite; 2-halogen-4-alkoxy pyridines are completely stable. Results are given of a quantitative study of the mobility of the halogen (determined by Volhard's method) and of the NO₂ group (conditionally, from the amount of NaOCH₃ consumed) in reactions with NaOCH₃ at 20°C and with boiling in alcohol and with NH(C₂H₅)₂ (with boiling in alcohol); these

Card 3/4

Substitution of a halogen and a ...

S/081/62/000/013/015/054
B158/B144

show that the 4-NO₂ group is considerably more reactive towards NaOCH₃ than 2-halogen in the pyridine ring. Of the halogens, in relation to NaOCH₃, the most mobile is Cl, the least mobile I; a difference in the mobility of Cl, Br and I is observed (at a relatively high reactivity) also with respect to NH(C₂H₅)₂; the most mobile is Br, the least mobile I.
8 references. [Abstracter's note: Complete translation.]

Card 4/4

TALIK, Z.

Substitution reactions of halogens and the nitro group in 2-halogen-4-nitropyridines. Bul chim PAN 9 no.9:567-569 '61.

1. Katedra Chemii Organicznej I., Politechnika, Wroclaw. Presented by T. Urbanski.

TALIK, Z.

The relative mobility of the halogen atoms in 2-halogenpyridines and
N-oxides of the 2-halogenpyridines. Bul chim PAN 9 no.9:571-574 '61.

1. Katedra Chemii Organicznej, Politechnika, Wroclaw. Presented by
T. Urbanski.

TALIK, Zofia; TALIK, Tadeusz

On the reaction of 2-halogen and 3-halogen-4-ni_xropyridine N-oxides with phosphortrihalogenides. Rocznik chemii 36 no.3:417-423 '62.

1. Katedra Chemii Organicznej I. Politechnika, Wrocław, i
Zakład Chemii Organicznej Katedry Chemii, Wyższa Szkoła
Ekonomiczna, Wrocław.

S/081/62/000/021/022/069
B141/B101

AUTHOR: Talik, Tadeusz, Talik, Zofia

TITLE: Preparation of some derivatives of 4-nitro-pyridine N-oxide

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 21, 1962, 173, abstract 21Zh163 (Roczn. chem., v. 36, no. 3, 1962, 539-544 [Pol.; summary in Ger.])

TEXT: The action of a mixture of $(\text{CH}_3\text{CO})_2\text{O}$ and H_2O_2 on pyridine, α -picoline, 2,6-lutidine, 2-chloro-, 2-bromo-, 2-iodo-, 3-chloro-, 3-bromo-, and 3-iodo pyridine (Ia-i), followed by nitration, yields the N-oxides of 4-nitro-I (IIa-i). The structure of IIg, i was confirmed by conversion with aniline to the N-oxide of 3-anilino-4-nitro-pyridine (III). 50 g Ia dissolved in 150 ml $(\text{CH}_3\text{CO})_2\text{O}$ 150 ml receives an addition of 30% H_2O_2 in portions, the solution is left to stand for 5-6 hrs at 20°C and is then kept for 30 hrs at $60-65^\circ\text{C}$, the excess of reagents is evaporated in vacuo at 60°C , the residue is dissolved in 50 ml concentrated H_2SO_4 and

Card 1/2

Preparation of some derivatives ...

S/081/62/000/021/022/069
B141/B101

added in portions to a mixture of 50 ml concentrated H_2SO_4 , 20 ml 20% fuming sulfuric acid, and 120 ml concentrated HNO_3 ($d\ 1.52$); the mixture is kept for 90 min at $100^\circ C$, cooled, and poured onto ice, neutralized to pH = 4-5 with solid $(NH_4)_2CO_3$, whereupon IIa is separated with concentrated NH_4OH , yield 63%, m.p. $162^\circ C$ (from water). Likewise substance II is obtained (yield in %, m.p. in $^\circ C$): b, 76.5, 155; c, 76.5, 156; d (here and below, 150 ml 30% H_2O_2 is added and the mixture heated for further 30 hrs) 69.4, 154; e, 62.5, 145; f, 61.5, 148; g ($C_5H_3ClN_2O_3$) (separated by extraction with $CHCl_3$), 64.5, 115; h, 84.2, 152; i ($C_5H_3IN_2O$) (obtained analogously to IIa), 56.4, 201. From the solution of 1 g IIg, 3 g aniline, and 20 ml alcohol which is boiled for 3 hrs and cooled, III, $C_{11}H_9N_3O_3$, is obtained, yield 0.7 g, m.p. $197-198^\circ C$ (from alcohol). Analogously, III is obtained from 1 g IIi and 3 g aniline in 40 ml alcohol by boiling for 6 hrs, yield 0.6 g. [Abstracter's note: Complete translation.]

Card 2/2

S/081/62/000/021/023/069
B141/B101

AUTHORS: Talik Zofia, Talik Tadeusz

TITLE: Effect of persulfuric acid on some 4-aminopyridine derivatives

PERIODICAL: Referativnyy zhurnal. Khimiya, no. 21, 1962, 173, abstract 21Zh164 (Roczn. chem., v. 36, no. 3, 1962, 545-548 [Pol.; summary in Ger.])

TEXT: When β -chloro-4-aminopyridine (I), β -bromo-4-aminopyridine (II), and 2-bromo-4-aminopyridine (III) are brought into reaction with $H_2S_2O_8$, the products are β -chloro-4-nitro-pyridine (IV), β -bromo-4-nitro-pyridine (V) and 2-bromo-4-nitro-pyridine (VI) respectively. Under analogous conditions, 2-chloro-, 2-iodo-, and β -ido-4-aminopyridine do not react with $H_2S_2O_8$. To 30 ml 20% fuming sulfuric acid 20 ml H_2O_2 is added dropwise under cooling by a mixture of ice and salt, then 2.5 g I dissolved in 7.5 ml concentrated H_2SO_4 is introduced at $0^\circ C$, the mixture is left to stand for 24 hrs at $20^\circ C$, poured onto ice, and neutralized with solid

Card 1/2

Effect of persulfuric acid on ...

S/081/62/000/021/023/069
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$(\text{NH}_4)_2\text{CO}_3$; the precipitate is dissolved in ether, and evaporation of the solvent produced the compound IV, $\text{C}_5\text{H}_3\text{ClN}_2\text{O}_2$, yield 71.2%, m.p. $25-26^\circ\text{C}$ (from benzene). Analogously V, $\text{C}_5\text{H}_3\text{BrN}_2\text{O}_2$, is obtained from 2.5 g II, yield 85.2%, m.p. $66-67^\circ\text{C}$. When 1 g III is brought into reaction with 5 ml H_2SO_4 and $\text{H}_2\text{S}_2\text{O}_8$ (from 6 ml H_2O_2 and 12 ml 20% fuming sulfuric acid), VI is obtained, yield 76.7%, m.p. 62°C (from alcohol). [Abstracter's note: Complete translation.]

Card 2/2

TALIK, Zofia

Research on the relative mobility of halogen atoms in 2-halogen pyridines and 2-halogen pyridine-N-oxides. Rocznik chemii 36 no.7/8:1183-1189 '62.

1. Katedra Chemii Organicznej I, Politechnika, Wroclaw.

TALIK, Zofia

Certain reactions of 2-halogen-4-nitropyridine. Rocznik Chemii 36
no. 9: 1313-1320 '62.

1. Katedra Chemii Organicznej I, Politechnika, Wrocław.

TALIK, Tadeusz; TALIK, Zofia

Reaction of 2-hydroxy-4-aminopyridine with nitrous acid. Roczniki
chemii 37 no.1:75-80 '63.

1. Zaklad Chemii Organicznej, Wyższa Szkoła Ekonomiczna,
Wrocław.

KHAUG, N.A. [Haug, N.]; kand.med.nauk (Tallin, ul. Roopa, d.19, kv.7);
TALIKHARM, A.A. [Taliharm, A.]

"Double block" following the use of lystenone. Vest. khir. 91
no.7 1966-67 JI '68
(MIRA 16:12)

I. P. Instituta eksperimental'noy i klinicheskoy meditsiny
AN Estonской SSR (dir.- doktor med. nauk P.A. Bogovskiy) i
Tallinskoy gorodskoy tuberkuleznoy bol'nitsy (glavnyy vrach
L.N.Osmirina).

TALIKOV, N.A.; BAKHOLDIN, S.V.; SERAVKIN, K.A.

Conveyors with a bushing-roller chain. Ferm. i spirt. prom. 30
no.1:32-33 '64. (MIRA 17:11)

1. Ryazanskiy likero-vodochnyy zavod.

TALIKOV, N.A.

Pathogenesis of pathological processes in the kidneys in pulmonary
tuberculosis. Sbor. trud. Uz. nauch.-issl. tub. inst. 3:126-190
'57. (MIRA 14:5)

(TUBERCULOUS) (KIDNEYS--DISEASES)

DEMUKHAMEDOV, A.A.; TALIKOV, N.A.

Renal functional and morphological changes in liver cirrhosis.
Med. zhurn. Uzb. no. 5-40-43 My'63
(MIRA 172)

1. Iz kafedry terapii (zav. - prof. A.S. Musikin) i kafedry
pathologicheskoy anatomii (zav. - prof. R.I. Denilova) Tash-
kentskogo instituta usovershenstvovaniya vrachev.

YANBAYEV, T.A.; TALIKOV, N.G.

Clinical anatomical comparisons of renal changes in suppurative processes in the body. Khirurgija 39 no.8:98-100 Ag '63.
(MIRA 17:6)

I. Iz khirurgicheskogo otdeleniya Tashkentskoy klinicheskoy bol'niцы nechtchnoy pomostchi (nauchnyy rukovoditel'- prof. S.A. Masumov; glavnyy vrach - zasluzhennyy vrach Uzbekskoy SSR T.Sh. Alimov).

TALIKOWA, Zofia

Chemical Abst.
Vol. 48 No. 3
Feb. 10, 1954
Organic Chemistry

(2) Chem
Chemistry of pyridine 1-oxide. Zofia Talikowa (Higher Polytech School, Wrocław, Poland). Wiedomosci Chem. 7, 160-88 (1953).—The following topics are reviewed: Substitution reactions of pyridine 1-oxide (I), chem. properties of substituted I, and reduction of the nitro deriv. of I with and without the cleavage of N-⁺ bond. 35 references.

Adam Sporzyński

8-31-53
djp

CICHOCKI, T.; RUTA, R.; TALIKOWSKA, H.

The distribution of some hydrolytic enzymes in the nephridium of the earthworm (*Lumbricus terrestris L.*). *Folia biol* 11 no.1: 69-83 '63.

1. Department of Histology, Medical Academy, Krakow. Head:
J. Ackermann, Ph.D.

~~FILM 1754730010-3~~, ~~REG. NO. 1000~~
GOLDSCHMIED, Aleksander; RAKAISKA, Zofia; TALIKOWSKI, Waclaw; BODZINSKA, Irena

Effect of intra-oral bile on the course of infectious jaundice. Polski
tygod. lek. 12 no.34:1319-1320 19 Aug 57.

1. (Z Oddzialu zoltaczek Kliniki Chorob Zakaznych A.M. w Lodzi; owczesny
kierownik Kliniki: prof. dr A. Goldschmied). Adres: Warszawa, ul.
Kasprzaka 17; III Zaklad Chorob Wewn. Instytutu Dosk. i Specj. Kadra Lek.
(BILE, therapeutic, hepatitis, infect., intra-oral admin. (Pol))
(HEPATITIS, INFECTIOUS, therapy,
bile, intraoral admin. (Pol))

TALIMAA, R.Yu.

Data from a study of specific and nonspecific reactivity in patients with tuberculosis of the lungs during combined treatment with antibacterial preparations and tuberculin. Probl.tub. 38 no.7:53-59 '60. (MIRA 14:1)

1. Iz Instituta tuberkuleza (dir. - chlen-korrespondent AMN SSSR N.A. Shmelev, nauchnyy rukovoditel' - doktor med.nauk F.L. Elinson) AMN SSSR.
(TUBERCULOSIS)

TALIMAA, R. Yu., Cand. Medic. Sci. (diss) "Changes of Allergical Reactivity in Patients with Tuberculosis of Lungs with Treatment Anti-bacterial Preparations and Tuberculin," Moscow, 1961, 23 pp. (Acad. Med. Sci. USSR) (KL Supp 12-61, 289).

TALIMAA R. J. med. m. kand.

An automatic spirograph without the water spirometer. Sveik.
Apsaug. no. 3:39-43 '64.

1. Lietuvos respublikinis tuberkuliozes mokslinio tyrimo
institutas.

TALIMAA R.^u, med.m.kand.

An automatic spirograph from a common water spirometer.
Sveik. apsaug. 9 no.3:39-43 Mr'64

1. Tuberkuliozes m.t. institutas.

TALIMOV I. A.

Identification of atypical dysenteric bacteria. Zhur.mikrobiol.
epid. i imun. 28 no.4:95-98 pp. 157. (M. 19-10)

.. In Tatarskogo gosudarstvennogo universiteta.

(ШИЗОТИЧНАЯ ДИСЕНТЕРИЯ)

identification of atypical strains by biochemical &
immunol. characteristics

TALMEYSTER, E.T.; RAUDSJK, T.A.

Study of the pathogenic properties of some freshly isolated
enteropathogenic Escherichia coli of varicous serological types.
Zhur. mikrobiol., epid. i immun. 42 no.7:21-24 Jl '65.
(MIRA 18:11)

1. Tartuskiy gosudarstvennyy universitet.

KUROL'SKII, A. I., and M. P. MEL'NIKOVA.

Vliyanie zakruchivaniia lopasti na aerodinamicheskie kharakteristiki avtozirova. Moskva, 1939. 34 p., diagrs. (TSAGI. Trudy, no. 396)

bibliography: p. 3h.

Title tr.: Effect of blade twist on the aerodynamic characteristics of an autogiro.

JA911.X65 no. 396

SO: Aeronautical Sciences and Aviation in the Soviet Union, Library of Congress, 1975.

TALIN, A.A.

Continuous action horizontal sterilizers. Kons. i ov. prom. 1⁴
no.10:23-24 O '59. (MIRA 12:12)

1.Izmail'skiy konservnyy kombinat.
(Sterilization)

TALIN, A. A.; SLYUSARENKO, N. A.

Production of grape juice at the Izmayl Cannery. Kons. i ov.prom.
15 no.10:6-8 0 '60. (MIRA 13:10)

1. Izmail'skiy konservnyy kombinat.
(Izmayl--Grape juice)

TALIN, A.O.

Aseptic preservation of food in the United States. Khar.prom.
no.2:91-94 Ap-Je '62. (MIRA 15:9)

1. Izmail'skiy konservnyy kombinat.
(United States—Canning and preserving)

BUTYAGINA, A. N.; VORONKOVA, O. I.; TALINSKAYA, A. F.; USHAKOVA, S. P.

Studying outbreaks of Boksin's disease in children's institutions.
Sov.med. 19 no.12:55-59 D '55. (MLR 10:9)

1. Iz Instituta virusologii AN SSSR i Moskovskogo chleastnogo
nauchno-issledovatel'skovo klinicheskogo instituta imeni M.F.
Vladimirevskogo
(HEPATITIS, EPIDEMIC)

ITSELIS, F. G.; YAMPOL'SKAYA, E. I.; ZALANZON, Ye. S.; MIL'NER, B. I.;
ROZENBAUM, G. I.; TALINSKAYA, A. F.

Focus of mixed diseases due to poliomyelitis and Coxsackie
[viruses] in a children's collective. Pediatriia no.6:15-19
'62. (MIRA 15:6)

1. Iz sanitarno-epidemiologicheskoy stantsii Moskvy i Instituta
pediatrii Ministerstva zdravookhraneniya RSFSR.

(POLIOMYELITIS) (COXSACKIE VIRUSES)

GRINEVICH, A.G.; TALIPOV, B.T.

Sensitivity of Streptococcus diacetilactis cultures to gamma rays. Uzb. biol. zhur. ? no.4:62-67 '65 (MIRA 11:84)

1. Institut botanikii AN USSR.

KARGIN, V.A.; SGOLOVA, T.I.; TALIPOV, G.Sh.

Plasticization of crystalline polymers. Part 1: Plasticization
of isotactic polystyrene and polyethylene terephthalate.
Vysokom.sod. 1 no.11:1670-1677 N '59. (MIRA 13:5)

1. Fiziko-khimicheskiy institut imeni L.Ya.Karpova.
(Styrene) (Terephthalic acid)

S/020/62/142/003/021/027
B101/B110

AUTHORS: Kargin, V. A., Academician, Sogolova, T. I., and Talipov, G.
Sh.

TITLE: Structure formation in crystalline polystyrene

PERIODICAL: Akademiya nauk SSSR. Doklady, v. 142, no. 3, 1962, 627-628

TEXT: The authors studied the formation of supermolecular structures in crystalline, isotactic polystyrene. The sample was fused onto an object glass, brought to test temperature (120, 145, 175°C) in a thermostat, and the structure formation was investigated and photographed in polarized light with 200 - 600fold magnification by means of an MIM-8 (MIN-8) microscope. The structure formation was found to be a complicated process, not ceasing with the formation of spherolites. Spherolites both grow and aggregate. At high temperature and prolonged crystallization, the number of growth centers is small, and regular spherolites are formed without disturbances. At low temperature, the growing rate is low but the number of centers is large. In this case, aggregation to bands takes place (length 25 - 500μ, width 1 - 120μ). Spherolites united to bands grow only
Carl 1/2

Structure formation in crystalline...

S/020/02/142/003/021/027
R101/3110

in width (possibly also in thickness) until laminae are formed and the growth ceases. The bands in one lamina are ordered, but their position in superimposed laminae does not coincide. The time of crystallization has the same effect on the formation of supramolecular structure as temperature has. Thus, systems of a desired structure can be produced. No molecules or molecule packages but spherulites of the order of magnitude of colloidal particles act as structural units in polymer systems. A similarity with the formation of gels and jellies is assumed. There are 4 references: 3 Soviet and 1 non-Soviet. The reference to the English-language publication reads as follows: F. Dannusso, G. Moraglio, J. Polymer Sci., 24, 161 (1957). ✓

ASSOCIATION: Fiziko-khimicheskiy institut im. L. Ya. Karpova (Physico-chemical Institute imeni L. Ya. Karpova)

SUBMITTED: October 11, 1961

Card 2/2

15.8100
S/020/62/142/004/05/022
B101/B110

15.8100
AUTHORS: Kargin, V.A., Academician, Sogolova, T.I. and Taliashvili, Sh.

TITLE: Structure formation in plasticized crystalline polystyrene

PERIODICAL: Akademiya nauk SSSR. Doklady, v. 142, no. 4, 1962, 844 p.

TEXT: Investigations of the supermolecular structure of nonplasticized polystyrene (DAN, 142, no. 3 (1962)) showed that secondary structures, bands or lamellas, composed of spherolites, developed between 110 and 215°C. The effect of plasticizers on structure formation and mechanical properties of polystyrene (PSt) was now investigated by the same method at 120, 145 and 175°C. The following data were found for the spherolite dimensions (in microns):

Duration of crystallization, min	Initial PSt	Concentration of plasticizer, % by volume					
		A		B		C	
		8	20	15	25	12	18
10	4	13	11	10	15	15	10
30	8	60	55	35	35	50	35
60	17	110	92	70	67	70	65
150	40	300	230	260	190	220	170

Card 17.2

Structure formation in...

S/02C/22/142/001/C-5/C22
B101/B110

A = cetyl chloride; B = dimethyl phthalate; C = dibutyl sebacate. Thus an increase in size of the spherolites occurred for all plasticizers, the formation of bands or lamellas, especially above 120°C, being suppressed. The effect of plasticizers was less marked at higher temperatures (175 - 215°C) since under such conditions the spherolites grow quickly growing even in nonplasticized PSt. The strength of plasticized PSt is increased with increasing plasticizer content (Fig. 4). X-ray pictures of initial and plasticized PSt showed only slight differences. Accordingly, the effect of plasticizer does not consist in a change of the primary crystalline structure (of the spherolite) of the polymer but in a suppression of secondary-structure formation. This also causes a reduction in strength which was investigated at a temperature 25°C higher than the vitrification temperature. There are 4 figures, 1 table, and 1 Soviet reference.

ASSOCIATION: Fiziko-khimicheskiy institut im L. Ya. Karpova (Physical chemical Institute imeni L. Ya. Karpova)

SUBMITTED: October 26, 1961

Card 2/4

S/0190/63/005/012/1809/1616

ACCESSION NR: APLC07979

AUTHORS: Kargin, V. A.; Sogolova, T. I.; Talipov, G. Sh.

TITLE: Supermolecular structure of plasticized and nonplasticized crystalline polystyrene

SOURCE: Vy'sokomolekulyarnye soyedineniya, v. 5, no. 12, 1963, 1809-1816

TOPIC TAGS: polymer, polystyrene crystalline, polystyrene, plasticized crystalline polystyrene, plasticization, plasticizer, secondary structure, glotule, spherolite, bundle, rod, supermolecule structure formation, plasticized polystyrene, mechanical property, polymer mechanical property, phthalic acid dimethyl ester, cetyl chloride, supermolecular structure

ABSTRACT: The supermolecular structure of plasticized (dimethyl-phthalate and cetyl chloride plasticizer) and nonplasticized crystalline polystyrene has been studied in temperature intervals from 110-215°C and at crystallization durations of 10 to 150 minutes. The structure-forming process was observed under a polarizing microscope MIN-3 with a 200-600 magnification. A new type of structuralization was discovered, in which spherolites serve as initial structure units, commensurable

Card 1/2

ACCESSION NR: AP4007979

in dimension with the colloidal particle size. It was found that the plasticizer has a substantial effect on the nature and dimensions of crystalline polymer secondary structure, and a definite correlation was established between plasticized polymer macrostructure and its mechanical properties. Orig. art. has: 16 micro-photographs and 1 chart.

ASSOCIATION: Fiziko-khimicheskiy institut im. L. Ya. Karpova (Physicochemical Institute)

SUBMITTED: 24Apr62

DATE ACQ: 20Jan64

ENCL: 00

SUB CODE: MA

NO REF Sov: 004

OTHER: 001

Card 2/2

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

PERMIK, I. (1910-1980)

Member, Central Committee of the Basque National Union, Basque People's Party, and the Basque People's Party, prevailing in Asturias. Also, member of the Basque People's Party, and the Basque People's Party, prevailing in Asturias.

b. Instituto Oficial de la Prensa

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

TALIPOV, N.A.; SKVALETSKIY, Ye.N.

Conditions of the Turgen' Hydroelectric Power Station from the
point of view of engineering geology. Izdatelstvo Nauk. Ser. est.
Izdat. nauk. 2 no.6, 143-148 '60. (IZ A 15.5)
(Kirghizistan. Hydroelectric power stations)

SUKHOVA, M.N.; ZAIROV, K.S.; GVOZDEVA, I.V.; ANDREYEVA, A.I.; NURULLAYEV,
D.Kh.; TALIQOV, M.Z.; MOSUNOV, V.B.; STOROZHEVA, Ye.M.; UMSNOVA,
A.M.; SHAMIRZAYEV, N.Yu.; AKMURZAYEV, T.A.

Fly control and its organization in Uzbekistan. Med.zhur.Uzb.
no.3:3-14 Mr '62. (MIRA 15:12)

1. Iz TSentral'nogo nauchno-issledovatel'skogo dezinfektsionnogo
instituta Ministerstva zdravookhraneniya SSSR (dir. - prof.
V.I.Vashkov) i sanitarno-epidemiologicheskoy organizatsii
Uzbekistana (glavnnyy gosudarstvennyy sanitarnyy inspektor-
kand.med.nauk K.S.Zairov).

(UZBEKISTAN--FLIES--EXTERMINATION)

KHAMRABAYEV, I.Kh.; TALIPOV, R.M.

Some results of biogeochemical (geobotanical) research in western
Uzbekistan. Uzb. geol. zhur. no.5:3-11 '60. (MIRA 13:11)

1. Institut geologii AN UzSSR.
(Uzbekistan--Ore deposits) (Phytogeography)

TALIPOV, R.M.

Characteristics of the concentration of nonferrous metals in the soils
and plants of the Sary-Cheku and Jen-Kalach regions (Uzbekistan). Geo-
khimia no.5:457-462 My '64. (MIRA 18:7)

I. Kh.M.Abdullaev Institute of Geology and Geophysics, Academy of
Sciences, Uzbek Soviet Socialist Republic.

TALIPOV, S.

Waters of the petroleum and gas fields of the Tajik depression.
Izv. vys. neoneb. zav., neft' i gaz 6 no. 11. 7-10 '63. (MIRA 17-9)

I. Moskovskiy institut neftekhimicheskoy i gazonoy promyshlennosti
im. akad. I.M.Gubkina.

TALIPOV, E.; SHIROKOV, V.Ya.

Waters and brines of the Mesozoic and Tertiary sediments of the
Tajik Depression in connection with their oil and gas potential.
Neftegaz.geol. i geofiz. no.12:12-13 '64.

(A121 18.3)

1. Moskovskiy ordena Trudovogo Krasnoznamennogo institut
naftekhimicheskoy i gazovoy promyshlennosti im. akad. Gittmana.

TALIPOV, S.; VAGIN, S.B.; SHUGRIN, V.P.

Gas content characteristics of waters of the Mesocenozoic sediments
in the southern Tajik depression. Izv. ys. ucheb. zav.; neft' i
gaz. 7 no.10:15-18 '64. (MIRA 18:2)

l. Moskovskiy institut neftekhimicheskoy i gazovoy promyshlennosti
im. akad. I.M. Gubkina.

TALIPOV, S.

Bromine and iodine content of underground waters in the Mubarek group of oil and gas fields. Dokl. AN Uz. SSR 21 no.9:44-46 '64.
(MIRA 19:1)

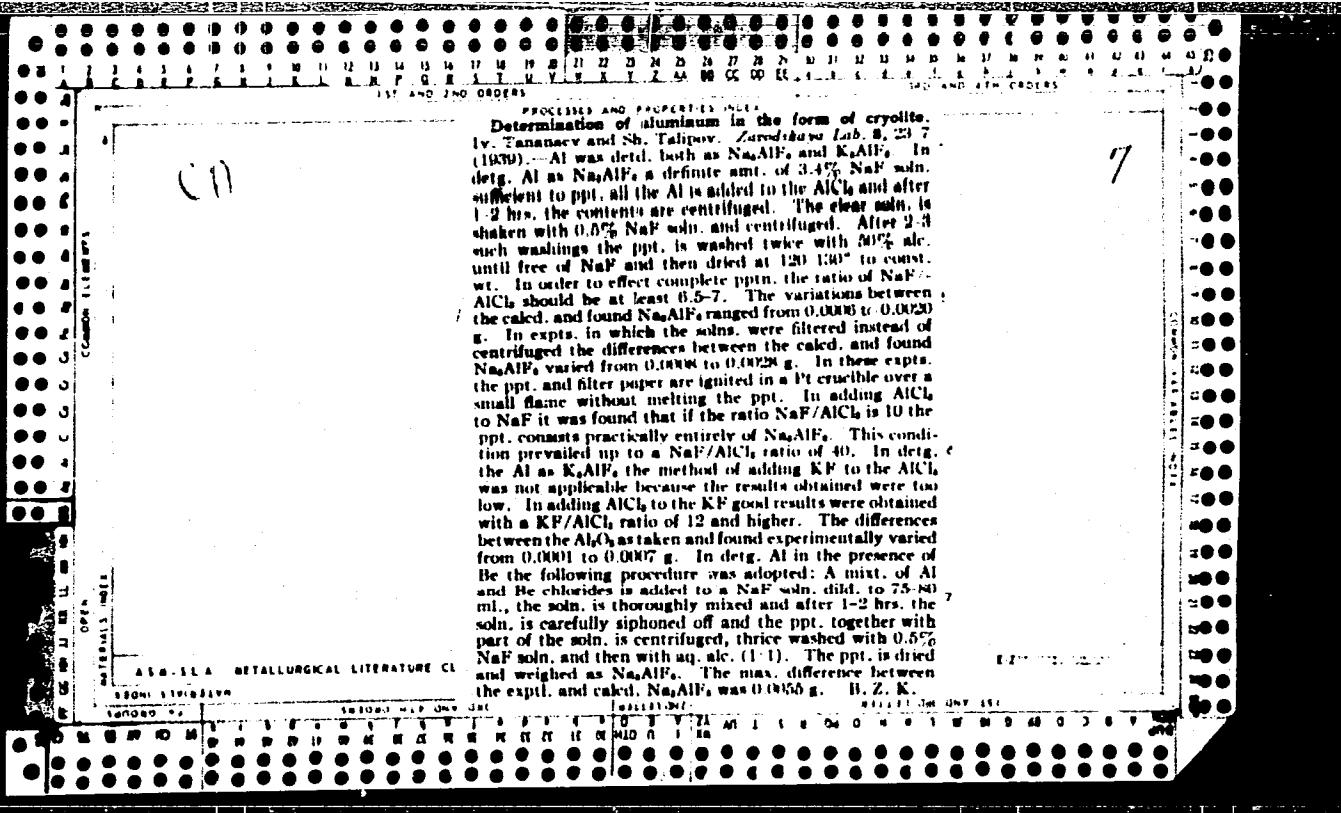
1. Institut geologii i razrabotki neftyanykh i gazovykh mestorozhdeniy Gosudarstvennogo geologicheskogo komiteta SSSR.

Use of fluorine compounds in chemical analysis. I. TANANAY AND S. TALIROV. (Bull. Acad. Sci. U.R.S.S., 1938, Ser. Chim., 547-553).—For the separation of Be from Ca and Mg a slight excess of NaF is sufficient; for separation from Al and Fe, 0.5 and 1.5 g. per l., respectively, is necessary.

L. J. J.

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"



Solubility of the double fluorides of aluminum and of the basic metals. I. Jananay and Sh. Talipov. *J. Gen. Chem. (U. S. S. R.)* 9, 1156 (1930).—An exptl method is given for the prepn. of Na_3AlF_6 , K_3AlF_6 and $(\text{NH}_4)_3\text{AlF}_6$, whose compns. correspond exactly to their formulas. The expts. were performed in a thermostat at 25° in vessels having Hg seals. Twelve hrs. was required for the satn. After the equil. had been reached a definite vol. was dried and weighed. The analysis of

the dry residue corresponded to the formula M_3AlF_6 . Expts. were also performed for the detn. of the solv. of Na_3AlF_6 and of K_3AlF_6 at different temps. At 0, 25, 50, 75 and 100° the solv. of Na_3AlF_6 in g./1000 g. of the soln. was 0.3184, 0.4155, 0.5032, 0.6002 and 0.6909 and that of K_3AlF_6 in 1000 g. of the soln. 0.8001, 1.1286, 2.6125, 3.4511 and 4.3828, resp. The solv. of $(\text{NH}_4)_3\text{AlF}_6$ at 0, 25, 50, 75, 100, 125 and 150° was 1.3128, 2.0853, 42.3622, 12.1217, 11.7031, 11.2310, 11.0421, 7.4106, 6.91. In contrast to the solv. of Na_3AlF_6 and K_3AlF_6 the solv. of $(\text{NH}_4)_3\text{AlF}_6$ does not rise steadily, but it has a max. at about 50°, after which it decreases sharply. At 10° and 25° the values for the solv. are the same. The unusual behavior of $(\text{NH}_4)_3\text{AlF}_6$ is evidently due to a change in its cryst. structure. A supposition that the salt dissolves incongruently was not verified, since the compn. of the dry residue was close to $(\text{NH}_4)_3\text{AlF}_6$. The difference in the solubilities of the 3 investigated salts led to an attempt to use a soln. of $(\text{NH}_4)_3\text{AlF}_6$ for the pptn. of the ions of Na and K. Preliminary expts. showed that both ions, used singly or together, were quantitatively ppd. as NaAlF_6 and KAlF_6 . Since the solv. of each ppt. was very small, attempts were made to det. the amt. of Na and K by microchem. means. Three graphs, 3 tables and 3 references are given.

W. R. Henn

Inst. of Metallurgy, Moscow, RSFSR

AVP-AKA METALLURGICAL LITERATURE CLASSIFICATION

CH

PROCESSES AND PROPERTIES OF...

Separation of titanium from aluminum and iron with fluorides. Sh. T. Talipov and Z. T. Sofelkova (Middle Asiatic State Univ.); "Zavodskaya Lab.", 13, 816-19 (1947).—It is proposed to utilize the differences in solv. for Na_2TiF_6 , Na_2AlF_6 , and Na_2FeF_6 in sepg. Ti from Al and Fe. To sep. Ti from Al, add to the neutral soln. of AlCl_3 and TiCl_3 3-4 times as much NaF as required to form AlF_6^- . Filter off the Na_2AlF_6 ppt., and wash with 0.5% NaF soln. and finally with EtOH. Filter and dry to const. wt. at 120-130°. To the filtrate add HCl or H_2SO_4 , evap. to about 50 ml., and ppt. the Ti with cupferron. Wash with 2% cupferron soln., ignite, and weigh as TiO_2 . Similarly, Fe can be sepd. from Ti with the Fe ptd. as Na_2FeF_6 . In the filtrate, the Ti can be detd. by adding H_2SO_4 , evap., to distil off the P as H_3SiP_3 , and finally titrating with ferric alum soln. B. Z. Kamish

TALIPOV, SH. F.

28936 Potenpiometricheskoy Titrovaniy Ftortdov. Soobshch, 2. Zavodskaya Laboratoriya,
1949, No. 9 S. 1031-34-Biblogr: 9 Nazv.

SO: Letopis' Zhurnal'nykh Statey, Vol. 39, Moskva, 1949

AMERI, SUL

Fluorine

Methods of quantitative determination of fluorine and the use of fluorine compounds for determining other elements. Sov. Sredniz. Akad., no. 27, 1947.

9. Monthly List of Russian Accessions, Library of Congress, ~~January 1952-1953~~, Uncl.

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

Malinov, V. M. et al. Khaderev, V. A. "Thermodynamic analysis of systems involving analytical literature. The system BaF₂-KF-H₂O at 25 degrees", Izvestiya Akad. Nauk SSSR, Chem., No. 4, p. 59-62, (March 1970), - Bibliography: 15 items.

SL: B-742, II March 1971, (Letter to English Statute, No. 10, 1971).

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

TAL'IOV, Sh. T.

3235?

TAL'IOV, Sh. T. i ANTIPOV, V. Ye. Ob Otkrytiu Malykh Kolichestv Zn, Ni, Co, Cu, Cd
i Mi v Prisutstvi Bol'shh Kolichestv Khroma Dopolny A. S. Nauk 1/3 CCC, 1949, № 9,
s. 16-17 Prezyurye Na Uzbek. Yaz-Bibliogr: s. 15

so: Istoricheskikh Churnal'nykh Statey Vol. 4

TALIPOV, Sh T.

enc.

(3)

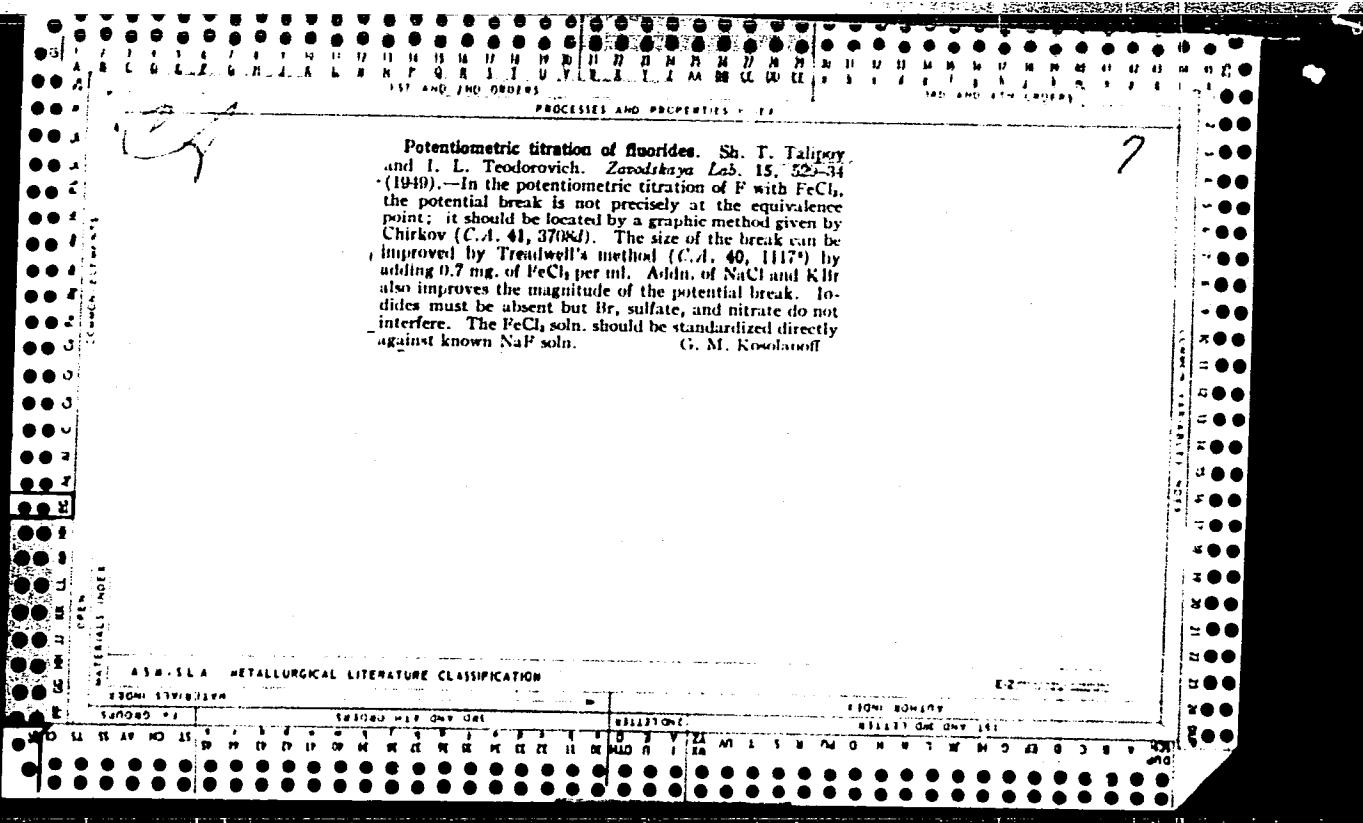
Chemical Abst.
Vol. 48 No. 4
Feb. 25, 1954
Inorganic Chemistry

New method of synthesis of chromium fluoride. Sh. T. Talipov and V. E. Antipov (Inst. Chem., Acad. Sci. Uzbek. S.S.R.). *Doklady Akad. Nauk Ussr.*, S.S.R. 1949, No. 12,

27-9 (in Russian).—Pure CrO₃ (40 g.), 240 g. 10% HF, 17 g. glucose, and 30 ml. EtOH, added in that order with agitation, until a spot test on filter paper shows green edges, gave CrF₃H₂O in the soln. The latter was concd. on a steam bath to a paste, stirred with 50 ml. warm H₂O, kept overnight, filtered through a Pt Büchner funnel, washed with EtOH and cold H₂O, yielding 85-90% of product after 2 recrystns. Purity of 99.60% was established by analysis.

G. M. Kosolapoff

MF
1-27-54



Potentiometric titration of fluorides. II. Sh. T. Talipov and I. L. Feodorovich. *Zavodskaya Lab.* 15, 1031-4(1949); cf. C.A. 43, 8942b. The best ratio of $\text{Fe}^{III}/\text{Fe}^{II}$ with the sharpest potential break is approx. 3.5; addn. of 0.1 ml. of soln. which is 0.033 M in FeCl_3 and 0.002 M in FeCl_2 per 10-20 ml. of test soln. gives about 700 mv./ml. potential break at the equivalence pt. in titrations of solns. which are about 0.2 M in fluorides. Typical titration curves are shown. G. M. K.

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

TALIPOV, Sh.T.

Certain general problems in the analytica chemistry of flourine.
Trudy SAGU no.15:67-84 '50. (MLRA 9:5)
(Fluorine)

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

TALIPOV, Sh.T.; KHADEYEV, V.A.

Physicochemical analysis of ternary aqueous solutions consisting of alkaline earth and magnesium fluorides and alkali metal fluorides. Ternary systems MgF_2 -- KF -- H_2O and MgF_2 -- NaF -- H_2O at 20°. Trudy SAGU no.15:85-100 '50. (MIRA 9:5)
(Fluorides) (Solution (Chemistry))

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Physicochemical analysis of ternary aqueous systems of fluorides of alkaline earth metals and alkali metals. I. Ternary systems $\text{BaF}_2\text{-KF-H}_2\text{O}$ and $\text{BaF}_2\text{-NaF-H}_2\text{O}$. Sh. T. Talipov and V. A. Khadeev. (Central Asian State Univ., Tashkent). *Zhur. Obrabotki Khim.* (J. Gen. Chem.) 20, 774-82 (1950); *J. Gen. Chem. U.S.S.R.* 20, 813-21 (Engl. translation). --The ternary systems were investigated by analyzing satd. solns. and wet residues according to the Schreinemakers method, all at 25°. No double salts or solid solns. are formed in either system. The solv. of BaF_2 in KF and in NaF was detd. at concns. of alkali fluoride up to 0.5 M. The exptl. results agree well with those predicted by the Debye-Hückel theory in its 2nd approximation. Values for KF concn. (M) and for the solv. of PbF_2 (M) are, resp.: 0.0, 9.21×10^{-2} ; 1.087×10^{-2} , 0.20×10^{-2} ; 5.011×10^{-3} , 1.60×10^{-3} ; 0.1001 , 7.5×10^{-4} ; 0.0020 , 2.2×10^{-4} . The values obtained for the solv. of PbF_2 in NaF are similar. II. Ternary systems $\text{SrF}_2\text{-KF-H}_2\text{O}$ and $\text{SrF}_2\text{-NaF-H}_2\text{O}$. *Ibid.* 783-8. --No double salts or solid solns. are formed in either system, the diagrams being qualitatively identical with those for the corresponding Ba compds. Solv. of SrF_2 was studied at concns. of NaF and KF up to 0.01 M. Values of NaF concn. (M) and SrF_2 solv. (M) are, resp.: 0.00, 0.02×10^{-2} ; 1.005×10^{-2} , 4.51×10^{-3} ; 3.000×10^{-3} , 2.00×10^{-3} ; 0.082×10^{-3} , 0.3×10^{-4} ; 1.000×10^{-4} , 5.0×10^{-5} . Data obtained with KF are similar. Arild J. Miller

TALIPOV, Sh. f.

Physico-chemical analysis of ternary aqueous systems containing
fluorides of alkaline earth metals and magnesium and iodides
of alkali metals. II. Ternary systems strontium fluoride-potassium
fluoride-water and strontium fluoride-sodium fluoride-water at 25°
S. N. Talyzin and S. A. Kostylev, Zh. Neorg. Khim., 1959,
Vol. 4, p. 88. U.S. Pat. 3,177,827. The properties at 25° are
described for the systems $\text{SrF}_2\text{-KHF}_2\text{-H}_2\text{O}$ and $\text{SrF}_2\text{-NaHF}_2\text{-H}_2\text{O}$.
No mixed crystals are formed. The solubility of SrF_2
in aq. KHF₂-H₂O is accurately represented by the Inaki-Hückel

formula in the vicinity of the solubility of SrF_2 in H₂O at
25° (see ref. 1). Omissions

TALIPOV, Sh.T.; ANTIPOV, V.Ye.

Investigation of solubility in the system chromium fluoride-sodium fluoride-water at 30°. Trudy Inst. Khim., Akad. Nauk Uzbek. S.S.R. 3, 206-13 '52.
(CA 47 no.22:11930 '53) (MLRA 6:3)

1. Uzbek Acad. Sci., Tashkent.

TALIPOV, Sh.T.; ANTIPOV, V.Ye.

Separation of large quantities of chromium from small quantities of zinc,
copper, nickel, cobalt, and cadmium. Trudy Inst. Khim., Akad. Nauk Uzbek.
S.S.R., Inst. Khim. 3, 214-22 '52. (MLRA 6:3)
(CA 47 no.22:12106 '53)

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

TALIPOV, Sh. T. OSEL'CHENKO, P. F.

"Solubility of Some Difficult to Dissolve Silver Salts in Solutions of the
Systems Ag_3PO_4 - K_2HPO_4 - H_2O and Ag_3PO_4 - K_2HPO_4 - H_2O at 25°

Dokl AN Uzb SSR, No. 5, 1952, pp 42-44

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W-31098, 26 Nov 54

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

TALIPOV, Sh.T.; ANTIPOV, V.Ye.

Investigation of solubility in the system aluminum fluoride-lithium fluoride-water at 25°. Trudy Inst. Khim., Akad. Nauk Uzbek. S.S.R. Inst. Khim. 3, 223-9 '52.
(MLRA 6:3)
(CA 47 no.22:11929 '53)

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

1A61PC4, 211, 1.

Determination of fluorine in insoluble fluorides by the method of I. V. Taranaev, Sh. T. Talipov, Z. T. Sofelkova, and T. B. Amirkhanova. *Trudy Sredneasiat. Gosudarst. Univ. (Tashkent)*, 33, No. 4, 75-82 (1952); cf. *C.A.* 47, 43.—When this earlier method was studied by detns. of CaF_2 , the abs. error was about 2%. The following modified method is proposed. Mix 0.1 g. of fluoride sample and 1 g. ferrosilicon in a 100-ml. round-bottomed flask connected with a thermometer, dropping funnel, and condenser; add 60 ml. of 12*N* H_2SO_4 through the funnel and then pass air into the mixt. while the flask is heated. The distd. H_2SiF_6 is collected in a flask contg. KCl. After distn. is complete (as shown by a temp. of 130° in the flask and by cooling of the delivery tube) any H_2SO_4 in the distillate is oxidized with I_2 and the color is removed with $\text{Na}_2\text{S}_2\text{O}_3$; the K-SiF_6 is neutralized to a methyl orange-indigo carmine end point, and then the distillate is titrated to a phenolphthalein end point. With this modified method the error averaged only about 0.30%; results of detns. on natural phosphates approximated those by the Th method. Errors tend to be neg., since the main source of error is incomplete distn. The changes in the method (use of ferrosilicon, passage of air, increasing the amt. of H_2SO_4 , etc.) improved the accuracy, shortened the distn. time, and simplified treatment of the distillate. F (11-65 mg.) was detd. as above.

Malcolm Anderson

R. G. S.

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TALIPOV, Sh. T.

✓ Potentiometric titration of microquantities of soluble fluorides. Sh. T. Talipov, I. L. Teodorovich, and N. P. Shestakova. Prudy i emtaziat. Gosudarstv. Univ. (Tashkent) 33, No. 4, 83-90 (1952). An aq. soln. of NaF was placed in a 5-ml. beaker, this was heated at 70-80° and a known vol. (an excess) of 0.0943M $\text{Ca}(\text{NO}_3)_2$ was added. After cooling, the soln. was dild. with an equal vol. of alc., and one drop of 0.01% $\text{K}_4\text{Fe}(\text{CN})_6$ soln. and an excess of solid NH_4Cl were added. The beaker was placed in an app. contg. a Pt electrode, which was also a stirrer, and a calomel reference electrode with a salt bridge. The soln. was stirred 1 min., let stand 1 min., and the e.m.f. was detd.; then the excess Ca^{++} was titrated with 0.500M $\text{K}_4\text{Fe}(\text{CN})_6$ according to the equation: $2\text{NH}_4^+ + \text{Ca}^{++} + [\text{Fe}(\text{CN})_6]^{4-} = \text{Ca}(\text{NH}_4)_2[\text{Fe}(\text{CN})_6]$. E (0.019-1.90 mg.) was detd. in this way. Addn. of 10 ml. of HOAc (to 0.380 mg. F) or of 2 mg. SiO_2 or 0.6 mg. Na_2SiO_4 with 2 drops of methyl red (to 0.051 mg. F) did not interfere.

Malcolm Anderson

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Talipov, Sh. T.

Determination of sodium as Na_2AlF_6 . Sh. T. Talipov and
Z. T. Solikova. *Izdat. Sredneaziat. Gosudarstv. Univ.*
(Tashkent) 33, No. 4, 91-93 (1952).— NaCl solns. of various
concn. (2-20 mg. as Na) were placed in tubes, $\frac{1}{4}$ their vol. of satd. $(\text{NH}_4)_2\text{AlF}_6$ (7.00 g./l. at 25°) was added, and the
solns. were centrifuged 4-5 min. The supernatant fluids
were decanted and the ppts. were washed 3-4 times with
0.5% $(\text{NH}_4)_2\text{AlF}_6$, then 2-3 times with 1:1 alc.- H_2O . The
ppts. were dried to 120-30° and weighed as Na_2AlF_6 (cryo-
lite). The method is rapid and comparable in accuracy to
the sulfate method, as far as on mirabilite by both methods
showed. P^- ions, given here by hydrolysis of $(\text{NH}_4)_2\text{AlF}_6$,
lowers the solv. of Na_2AlF_6 ; the presence of 0.1 mole/l. of
 P^- reduces the solv. of Na_2AlF_6 from 1.86×10^{-3} (in pure
 H_2O) to 6.3×10^{-3} moles/l. Presumably the presence of
other alkali metals interferes with this method. M. A.

HALIPOV, S. M.

✓ Gravimetric determination of magnesium as KMgF_3 .
Sh. T. Halipov and Z. T. Sofelkova. *Trudy Sredneasiat.*
Zhurnalist. Upr., 33, No. 4, 85-88 (1952).—Solns. contg. var.
ious concns. of MgSO_4 were placed in tubes with 10 ml. of
6.5-7% KF soln., and the colloidal ppts. that formed were
centrifuged. The supernatant liquids were drained off and
the ppts. were washed 5 times with 60% alc. (aq.) and 3
times with 10% alc. the ppts. being centrifuged after each
washing. The ppts. were dried at 130° , ignited to const.
wt. at $450-500^\circ$, and weighed as KMgF_3 . To det. Mg in
Al alloys, dissolve a 2-g. sample of alloy in 35 ml. of 25%
NaOH and filter; wash the ppt. with hot 1% Na_2CO_3 ; and
then redissolve in 40 ml. of hot HCl contg. some HNO_3 .
Neutralize the combined filtrates with NH_4OH , and remove
Cu, Pb, Fe, and other elements with H_2S ; then det. Mg as
above and also by the phosphate method. The method is
recommended for up to 0.1 g. MgO . Best results are given
with $\text{Mg}^{++}:\text{KF}$ molar ratio of 1:3; when this ratio is too
high, MgF_2 may be formed. *Malcolm Anderson*

TALIPOV, Sh.T.

USSR.

Solubility of some slightly soluble silver salts in electrolyte solutions of the systems: $\text{Ag}_2\text{O}-\text{Na}_2\text{HPO}_4-\text{H}_2\text{O}$ and $\text{Ag}_2\text{O}-\text{K}_2\text{HPO}_4-\text{H}_2\text{O}$ at 25°. Sh. T. Talipov and P. F. Obel'shchenko. Doklady Akad. Nauk Uzbek. SSR. 1953, No. 5, 42-47. Referat. Zhur. Khim. 1954, No. 17845. -- The solv. of Ag_2PO_4 was studied in solns. of Na_2HPO_4 and K_2HPO_4 in concns. of 10^{-2}M to satn. The solid phase at all points of the studied system corresponded to pure Ag_2PO_4 . The solv. of Ag_2PO_4 did not change on increasing the concn. of the added salt to 10^{-1}M but increased at higher concns. The solv. product of Ag_2PO_4 increased from 1.89×10^{-8} to 1.21×10^{-9} in 0.826M K_2HPO_4 and to 1.17×10^{-12} in 1.41M K_2HPO_4 . At the same time the activity product practically did not change. The activity coeff. was calc'd. with the aid of the 2nd approximation of the Deby-Hückel equation.

M. H. [initials]

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

TALIPOV, SH. T.

"Solubility of Some Difficult-to-Dissolve Silver Salts in Solutions of
Electrolytes" Dokl AN Uzb SSR, No 7, 1953 pp 29-31

Talipov, Sh. T. ; Obel'chenko, P.F.

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W-31098, 26 Nov 54

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

"Determination of Fluoride Ions Using the Ionometric Method," J. Am. Chem. Soc., No. 5, 1933, 55, 22-31.

Investigated the accuracy of the ionometric titration of gallium bromide solution using 1% malachite green indicator. The error was $\pm 1.6\%$. (J. Am. Chem. Soc., 1933)

Re: Cum No. 40, 16 Aug 55

"APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3

TALIPOV, Sh.T.; AKHMEDOV, K.S.

In memoriam Boris Georgievich Zaprometov. Kolloid. Zhur. 15,223-4
(MLRA 6:5)
'53,
(CA 47 no.17:8428 '53)

APPROVED FOR RELEASE: 07/13/2001

CIA-RDP86-00513R001754730010-3"

Talipov, Sh.T.

in memoriam Boris Georgievich Zanrometov. Sh. T.
Talinov and K. S. Akhmedov. Colloid J. U.S.S.R. 15,
226-30 (1953) (Engl. translation).—See C.A. 47, 8428a.
H. L. H.